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162 CHAPTER 6: THERMOCHEMISTRY To Convert The Answer To Joules, We Write: $101.3 \text{ J} \cdot 0.18 \text{ L} \cdot \text{Atm} = 18.2 \text{ J}$
 $1 \text{ L} \cdot \text{Atm} = 101.3 \text{ J}$
 -18 J 6.17 An Expansion Implies An Increase In Volume, Therefore W Must Be -325 J (see The Defining Equation For Pressure-volume Work.) If The System Absorbs Heat, Q Must Be $+127 \text{ J}$. The Change In Energy (internal 2th, 2024

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To Thesum Ofenergyin Form Heat Exchanged Between System And Surroundings (q) And The Work Done On The System (w). $\Delta E_{\text{system}} = Q + W$ But The Laws Of Thermodynamics Are Far More Profound Than Simply Describing The Efficiency Of Steam Engines. T 1th, 2024

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