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CHAPTER 12 GASES AND KINETIC-MOLECULAR THEORY

3 Boyle's Law: The Volume-Pressure Relationship {V $\propto 1/P$ Or{V= K (1/P) Or PV = K{P1V 1 = K 1 For One Sample Of A Gas. {P2V 2 = K 2 For A Second Sample Of A Gas. {k1 = K 2 For The Same Sample Of A Gas At The Same T. {Thus We Can Write Boyle's Law Mathematically As P 1V 1 = P 2V 2 Feb 24th, 2024

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