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Mole Fraction Sugar Of Solution $\chi_{\text{Sugar}} = (0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$ Mole Fraction Water Of Solution $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$ (or $1 - \chi_{\text{Sugar}}$) Molality $M = (.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$ Molarity $M = (.292 \text{ Mol}) / (1$

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Calculating Molality From Molarity

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1. A Solution Is Prepared By Mixing 100.0 G Of Water, H_2O , And 100.0 G Of Ethanol, $\text{C}_2\text{H}_5\text{OH}$. Determine The Mole Fractions Of Each Substance. 2. The Molality Of An Aqueous Solution Of Sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3. Calculate The Molality Of 25.0 Grams Of May 6th, 2024

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