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Equation That Represents The Oxidation Reaction That
Occurs Is A) $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$ B) $4 \text{HCl} + \text{MnO}_2$
 $\rightarrow \text{MnCl}_2 + 2 \text{H}_2\text{O} + \text{Cl}_2$ C) $2 \text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 +$
 $\text{H}_2\text{O} + \text{CO}_2$ D) $2 \text{HCl} + \text{FeS} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$ 21.Which
Equation Represents An Oxidation-reduction Reaction?

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14. Questions 14-17 The Spontaneous Reaction That
Occurs When The Cell In The Picture Operates Is As
Follows: $2\text{Ag}^+ + \text{Cd}(\text{s}) \rightarrow 2\text{Ag}(\text{s}) + \text{Cd}^{2+}$ (A) Voltage
Increases. (B) Voltage Decreases But Remains $> 5\text{V}$,
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