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2024CHEMISTRY 9701/01 1 Hour10 When Vanadium(II)  
Compounds Are Dissolved In Water, The Following  
Equilibrium Is Established.  $V^{2+} + H_2O \rightleftharpoons V^{3+} + H_2O + OH^-$   
What Would Alter The Composition Of The  
Equilibrium Mixture In Favour Of The  $V^{2+}$  Ions? A  
Adding An Acid B Adding A Reagent That Selectively  
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Equation Representing The Lattice Energy Of AgBr. (ii)  
Use The Following Data To Calculate A Value For The  
Lattice Energy Of AgBr(s). First Ionisation Energy Of  
Silver = +731kJmol<sup>-1</sup> Electron Affinity Of Bromine =  
325kJmol<sup>-1</sup> Enthalpy Change Of Atomisation Of Silver  
= +285kJmol<sup>-1</sup> Enthalpy Change Of Atomisation Of  
Bromine = +112kJmol<sup>-1</sup> ... 4th, 2024.

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Without Replacement, Is Equal To: MCQ 11.75 If X  
Represent The Number Of Units Having The Specified  
Characteristic And N Is The Size Of The Sample, Then

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