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### **Dr.M.T Bhoite1, Swapnil Gavhane, Mitali Gore, Sukanya ...**

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Example 9-4. Absorption Of SO<sub>2</sub> The Molar Density Of The Feed Gas May Be Determined From The Formula

$(\text{mol/L}) = 12.2P(\text{atm})T^{-1}(\text{K})$  Where P And T Are The Absolute Pressure And Temperature Of The Gas. The Following Data Are Taken For The Feed Gas: T=750F, P=150 Psig, H=210 Mm, R=82.4 And For The Outlet Gas, R=11.6. Determine The (a) Mass Fraction Of SO 11th, 2024

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Elements. For Example, Sodium Is A Shiny Metal, And Chlorine Is An Irritating Gas. But The Compound Formed From These Two Elements, Sodium Chloride, Shows Different Characteristics As It Is A Crystalline Solid, Vital For Biological Functions. Physical Classification Solid Mixtures Ice Water Water Vapour Homogeneous 10th, 2024

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$7 + 6 \text{ KI} + 7 \text{ H}_2\text{SO}_4 \rightarrow \text{Cr}_2(\text{SO}_4)_3 + 4 \text{ K}_2\text{SO}_4 + 3 \text{ I}_2 + 7 \text{ H}_2\text{O}$  A) How Many Moles Of Potassium Dichromate ( $\text{K}_2\text{Cr}_2\text{O}_7$ ) Are Required ... = 407.9 g AgBr This Is The Theoretical Yield Yields 22 B) Calculate The Percent Yield If 375.0 g Of Silver Bromide Was Obtained From The Reaction Theoretical Yield = 407.9 g AgBr Percent Yield =  $100 \times \frac{\text{Actual Yield}}{\text{Theoretical Yield}}$  20th, 2024

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