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=  $k[D]^1 X[B]^1 Y$   $k[D]^2 X[B]^2 Y$  Equation 5 In This Equation  $k$  Cancels Out.  $[D]^1 = [D]^2$  Because We Are Starting Off With The Same Initial Concentration Of  $A$  In Each Trial. The Order Of Species  $D$  Which Is  $X$  Is Also Constant. Canceling Terms We Have Left:  $Rate_1 = Rate_2 = k[B]^1 Y$   $k[B]^2 Y$  Equation 6 Consider The Following Chemical Reaction: Jul 2th, 2024

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Laws PLAN: We Inspect The Exponents In The Rate Law, Not The Coefficients Of The Balanced Equation, To Find The Individual Orders. We Add The Individual Orders To Get The Overall Reaction Order. (a) The Exponent Of  $[\text{NO}]$  Is 2 And The Exponent Of  $[\text{O}_2]$  Is 1, So The Reaction Is Second May 2th, 2024

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2and Cl 2to FormHCl. 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For A Reaction  $H_2 + Cl_2 \rightarrow 2HCl$  The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As  $X \rightarrow Y$  Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly Feb 2th, 2024



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B. Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds  
Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules  
KE  $E > E_A$ , Activation Energy - Energy Barrier That Molecules Have To Surmount In  
Order To React. Energy Is Needed To Break Reactant Bonds (endothermic Process).

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