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3301/2H (Paper 2 ... $(2 \times 1.67 \times 10^{-24}) + (2.66 \times 10^{-23}) = 2.994 \times 10^{-23} = 2.99 \times 10^{-23}$ (I Have Given The Answer To The Same Degree Of Accuracy As The Data We Were Given) B) The Above Is The Mass Of 1 Molecule Of Water. To Find Out How Many Molecules There Are In 1 Gram We Divide 1 By The Mass Of 1 Molecule. $1 \div (2.994 \times 10^{-23}) = 3.34 \times 10^{22}$ Question 15 Feb 3th, 2024.

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